

Azadirachta indica (Neem) fruit latex based iron oxide magnetic nanoparticles (FeO-MNPs): Synthesis and green treatment of reactive yellow RY-176 dye

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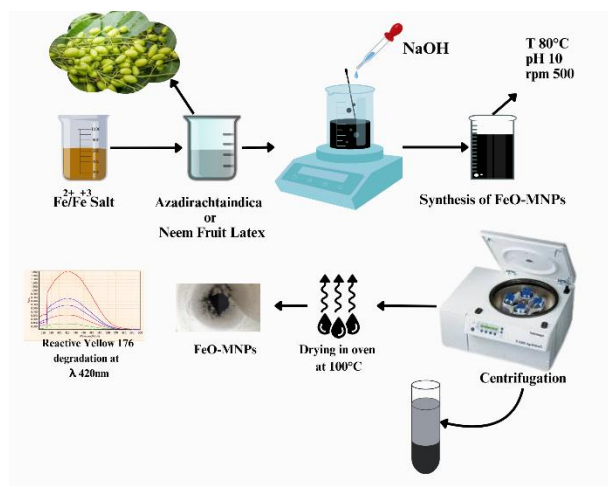
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Graphical abstract



Abstract

A considerable amount of wastewater is discharged by the textile industry primarily containing various pollutants, including dyes. The Reactive Yellow 176 dye is particularly notorious for its persistence and adverse effects on ecosystems. In response, the current work focuses on the preparation and application of iron oxide magnetic nanoparticles (FeO-MNPs), using unripe *Azadirachta indica* or Neem Fruit Latex (NFL) for the effective dye removal from aqueous media through the phenomena of adsorption. Iron oxide magnetic nanoparticles were synthesized using green co-precipitation method at mild temperature, employing *Azadirachta indica* fruit latex and ferric / ferrous chloride precursors. Various techniques of microscopy, spectroscopy, crystallographic and particle sizing were used to determine the synthesis of biogenic FeO-MNPs and for its characterization. Synthesized nanoparticles were in spherical shape with an average crystalline size of 16.5 nm, confirmed by X-ray Diffraction. The adjustability of the core size of particles, based on the latex concentration was confirmed by Atomic Force Microscopy whereas, Fourier Transform Infrared

Spectroscopy demonstrated the presence of different functional groups that worked as reducing agents and contributed for stabilization of nanoparticles. These NFL based particles exhibited high adsorption capacity and rapid removal kinetics against reactive yellow 176 dye. Influence of dye concentration, pH and contact duration were thoroughly investigated, highlighting conditions that maximize removal capacity. For a dye concentration of 30 mg L⁻¹, 99.95 % removal efficiency was achieved at pH 4 with removal capacities of 29.89 mg g⁻¹ in 250 minutes and retained 78.52 % of their adsorption efficiency after three consecutive reuse cycles. The adsorption process was best described by pseudo-second-order (PSO) kinetic model, indicating that chemical interactions played a significant role in dye removal. Overall, this study sheds light on the potential applications of biogenic iron oxide nanoparticles in industrial effluent treatment, providing an effective, efficient and ecofriendly approach to mitigate aquatic pollution.

Keywords: Reactive Yellow RY-176 dye; Iron oxide magnetic nanoparticles; Kinetic Models; Industrial Effluent Treatment and *Azadirachta indica* or Neem Fruit Latex (NFL).

1. Introduction

Contaminated water is the key ecological concern that is being faced all over the world induced by rapid industrialization and urbanization. This is largely due to the discharge of pharmaceutical active compounds and synthetic dyes in the industry's effluents (Sayed *et al.*, 2024). Although the textile industry plays a vital role in driving global economic progress, it also carries the burden of contributing substantially to water pollution (Maruthai *et al.*, 2025; Fernandes *et al.*, 2019; Periyasamy, 2024). Among these, reactive yellow (RY) 176 dye stands out for its resilient nature and adverse environmental impact on the ecological system. It is a synthetic dye commonly used in textiles, inks and cosmetics. RY 176 is known for its vibrant yellow color and is frequently

established. This type of study is crucial for optimizing and designing efficient adsorption systems for the pollutants removal from industrial effluents before their discharged into the aquatic environment (Al Aboody, 2019^b).

2. Methodology

2.1. Materials

Unripe green fruit of *Azadirachta indica* or neem were picked in the morning from University of Karachi, Pakistan, Ferric (III) chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$), Ferric (II) chloride tetrahydrate ($\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$) was used as Fe precursor, anionic reactive yellow dye ($\text{C}_{29}\text{H}_{21}\text{ClN}_8\text{Na}_4\text{O}_{16}\text{S}_5$), Deionized distilled water (DDW), sodium hydroxide (NaOH) and ethanol were purchased from Sigma Aldrich.

2.2. Preparation of Plant Material

Azadirachta indica unripe fruits were thoroughly rinsed with distilled water. Milky white latex as shown in **Figure 2A** was collected drop by drop in a sterile glass container by gentle pressing. The solution was filtered twice with filter paper and stored at 4°C for subsequent phytochemical analysis. Crude latex of 3 mL gently mixed with 100 mL DDW to prepare 3% latex solution. The solution was then used as plant material for further process.



Figure 2. (a) Milky white latex of unripe fruit of *Azadirachta indica*; (b) Visible color changes from orange to brownish black with precipitates

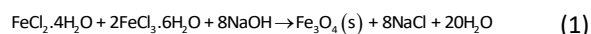
2.3. Phytochemical Analysis of Neem Fruit Latex (NFL)

Phytochemical screening of *Azadirachta indica* or neem fruit latex was performed to explore the availability of secondary metabolites following the methods detailed by (Ahmed *et al.*, 2018; Vinay *et al.*, 2018; Kenneth *et al.*, 2017).

2.4. Biogenic Synthesis of Iron Oxide Magnetic Nanoparticles (Adsorbent)

The biogenic synthesis of magnetic iron oxide nanoparticles was carried out based on the methods reported by Danjuma *et al.*, 2022; Arsalani *et al.*, 2018) with certain modifications. Firstly, 13.5 g and 4.90 g of ferric (III) and (II) respectively, were dissolved separately in deionized distilled water at room temperature. Fe^{3+} solution was mixed with Fe^{2+} to achieve $\text{Fe}^{3+}/\text{Fe}^{2+}$ molar ratio of 2:1. 50 mL of 3% neem fruit latex was added drop wise into 250ml of ferric chloride precursor and the mixture was agitated at 80°C using a magnetic stirrer set at 500 rpm. The gradual addition of freshly prepared 2 M NaOH adjusted the pH to 10, leading to a color change from orange to brownish black with the precipitates

formation. This transformation indicated the reduction of ferric chloride and the successful synthesis of iron oxide nanoparticles similar results reported by (Kanagasubbulakshmi & Kadirvelu, 2017^b; Mohandoss *et al.*, 2022). After resting at room temperature for 10 - 15 minutes, centrifuged at 4000 rpm for 15 minutes as part of the purification step. The purification process was repeated three times, including washing with ethanol and then dried at 100°C in a china dish resulting in black ferric oxide nanoparticles. The reactions involved are presented in equations 1 and 2.



2.5. Characterization of FeO-MNPs

Instruments used for characterization includes X-ray Diffractometer at 40 kV, 25 mA at 25°C in the range from 2 theta = 5.0° to 80.0°, FTIR analysis by Bruker, Vector-22, Spectrophotometer in the range of 4000-400 cm^{-1} , Scanning Electron Microscopy 450 (dwell 10 μs) Thermo Fisher Scientific armed with Energy Dispersive X-ray spectroscopy, Atomic Force Microscopy (AFM - 5500 Agilent Technologies) and UV-vis spectrophotometer Model-UV Thermo Scientific, (Evolution 300).

2.6. Adsorbate Reactive Yellow 176 Dye Adsorption Activity

A 30 mg L^{-1} standard stock solution of RY176 was prepared in DDW. The initial pH of the dye solution was adjusted to 4, 7 and 9 and contact times ranging from 0 to 250 minutes were tested using a constant dose of FeO-MNPs (5 mg) adsorbent. The solution's pH was adjusted by using 0.1 M HCL and NaOH and change in color intensity was calculated using a UV-Vis Spectrophotometer at 420 nm. The RY176 dye removal efficiency was calculated using equation 3. Three replicates were performed for adsorption studies.

$$\text{Adsorption Efficiency}(\%) = \frac{C_i - C_e}{C_i} \times 100\% \quad (3)$$

Where C_i and C_e represent the concentration of initial and equilibrium phase of RY 176 (mg L^{-1}) respectively. The amount of adsorbate (dye) adsorbed onto the adsorbent (iron nanoparticles) can be determined by using mass balance relationship as expressed by the equation 4.

$$\text{Adsorption Capacity}(\text{mg/g}), q_e = \frac{C_i - C_e}{W} \times V \quad (4)$$

Where W denotes adsorbent weight in gram, and V corresponds to the solution volume in liters.

2.7. Adsorption Kinetic Model

The dye RY176 removal potential was assessed, and the adsorption isotherm was determined employing Langmuir, Freundlich and sips isothermal models. The equilibrium data were subsequently evaluated using kinetic models, specifically pseudo-first order (PFO) and pseudo-second order (PSO). The most common models used to study the kinetics of adsorption (Khatun *et al.* 2022). The correlation coefficient (R^2) is considered as a

measure of consistence between proposed models and experimental data for current study.

2.8. Reusability Analysis of FeO-MNPs

Reusability assays of synthesized NFL based FeO-MNPs were performed to assess their commercial applicability. Following the adsorption process, the nanoparticles rinsed with deionized water and dried at 70°C to achieve desorption. The recovered nanoparticles were then reused as adsorbent, and the procedure was repeated for three consecutive cycles by maintaining other parameters i.e. dose of adsorbent, concentration of adsorbate, pH, and contact time throughout the experiments.

3. Result and Discussion

3.1. Screening of Phytochemical

Unripe neem fruit latex contains various bioactive compounds as listed in **Table 1**, which actively contribute to the formation of iron nanoparticles. Phenolic compounds, flavonoids, and tannins contain hydroxyl (-OH) and carboxyl (-COOH) groups, which can act as reducing agent by donating electrons to Fe^{3+} ions facilitating to their reduction into Fe^{2+} ions, a key step in Fe_3O_4 (magnetite) nanoparticle formation. These functional groups also act as capping agents preventing agglomeration and influencing their dispersion in aqueous solution. Similarly, aromatic rings present in flavonoids and other polyphenols can engage in π - π stacking interactions with iron oxide surface, enhancing nanoparticle stability.

Table 1. Screening of Phytochemicals in *Azadirachta indica* or Neem Fruit Latex (NFL)

Phytochemicals Components	Test	Interferences
Alkaloids	Dragendroff's	+++
Flavonoids	Lead acetate	+++
Saponine (Frothing)	Frothing	++
Phenolic Compounds	Ferric Chloride	+++
Tannins	Braymer's	++
Carbohydrates	Molish's	+
Reducing Sugars	Benedict	++
Glycosides (Cyanogenic)	Keller-Killani Test	++

3.2. Visual observation of FeO-MNPs

The color of the solution (iron oxide nanoparticles with latex) changed from orange to brownish black with precipitate formation at the bottom as clearly observed in **Figure (2B)** indicating reduction in metal ions. According to Jayandran *et al.*, 2015 dark brown color clearly indicates the formation of well reduced and stabilized iron oxide nanoparticles. The optimal conditions for synthesizing iron oxide magnetic nanoparticles were found to be pH 10, temperature of 80°C and reaction time of 45-60 minutes. Furthermore, the resulting nanoparticles exhibit magnetic property in the aqueous solution attracted by external magnet field and deposited at the side of the magnetic rod as shown in **Figure (3A)**

and **(3B)**. This phenomenon was observable only in the presence of applied magnetic fields suggestive of the formation of magnetic nanoparticles with a strong recommendation for recollect and reuse easily for the removal of dye. Notably, in the control experiment without neem latex, no color change and precipitate formation were observed, highlighting its essential role in the synthesis process.

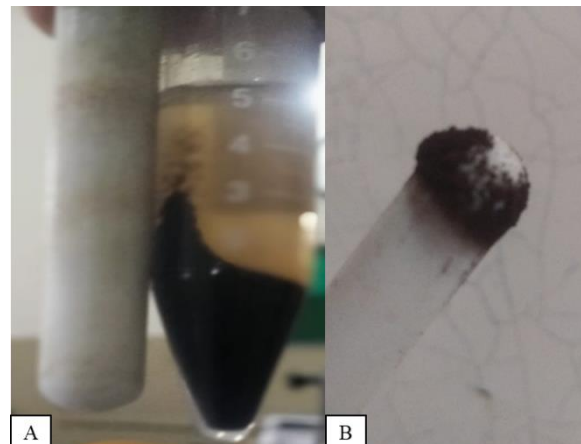


Figure 3. (a) Movement of FeO-MNPs due to external magnetic field; (b) FeO-MNPs attracted at the side of magnetic rod

3.3. Assay through UV-Vis Spectrophotometry

UV visible spectroscopy is a critical tool for initial characterization of nanoparticles. Typically, iron oxide nanoparticles show absorption band around 280-420 nm as mentioned by Tyagi *et al.*, 2021; Pindiga *et al.*, 2023. UV-visible spectral analysis, as shown in **Figure (4)** displayed an intense absorbance at 324 nm with a band gap value of 3.27 eV ascribed to the excitation of surface plasmon resonance in the FeO-MNPs. High peak intensity is attributed to strong interaction between light and the oscillating surface electrons indicating photosensitivity of the synthesized particles as discussed by Pattanayak & Nayak, 2013).

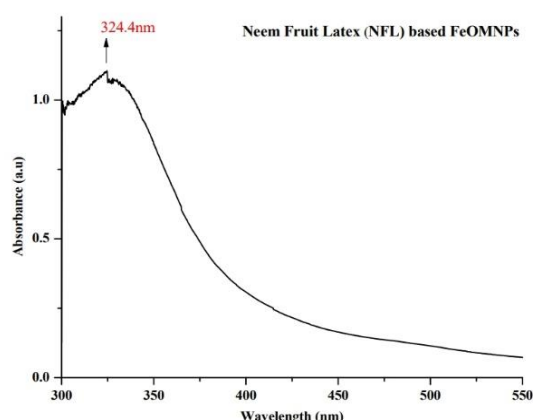


Figure 4. UV Spectrum of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

3.4. Assay through FT-IR Spectroscopy

IR spectrum of the NFL demonstrated the presence of various functional groups through characteristic absorption bands, which were essential for the reduction, capping and stabilization as shown in **Figure (5A)** including hydroxyl group of phenols, -CH bond of alkyl, alkenyl and

aryl groups, NH group of amines and amides, carbonyl (C=O) groups of aldehyde, ketones and amides, aryl C=C groups and C-O, C-N bonds of different organic molecules. FTIR spectra of FeO-MNPs observed various peaks within the range of 400-4000 cm^{-1} as depicts in **Figure (5B)** and confirmed the reducing role of neem fruit latex in the formation of FeO-MNPs. The peak at 3746.1 cm^{-1} corresponds to the O – H stretching (free hydroxyl) due to moisture or phenolic groups whereas, broad O – H and N – H stretching denotes the hydrogen bonding from flavonoids or amines present in latex. C $\equiv$ C or CO₂ asymmetric stretching at 2359.0 cm^{-1} indicates possible absorbed CO₂ or unsaturated groups and weak overtone relate to conjugated system at 2005.1 cm^{-1} . The strong peak at 1623.4 cm^{-1} C = O or C = C stretching (aromatic band) from flavonoids or tannins present in latex. O – H bending or COO⁻ symmetric stretch at 1419.3 cm^{-1} indicative of carboxylic acids or esters whereas 1076.7 cm^{-1} is the C – O – C or C – O stretching confirms presence of polysaccharides or phenolic groups. Peak at 767.6 cm^{-1} represents out-of-plane C – H bending related to aromatic compounds whereas, peaks at 538.7 cm^{-1} and 511.3 cm^{-1} are characteristic of metal-oxygen bonds (Fe – O) vibrations in the region of 550-500 cm^{-1} confirm formation of iron oxide nanoparticles (Sudhakar et al., 2022). There was significant loss of IR absorption peaks in FeO-MNPs spectrum which suggests successful nanoparticle formation. The biomolecules act as reducing and stabilizing agents by donating electrons to reduce metal ions to their metallic states and often attach to the nanoparticle surface through coordination bond which causes their IR absorption peak to either weaken or disappear altogether.

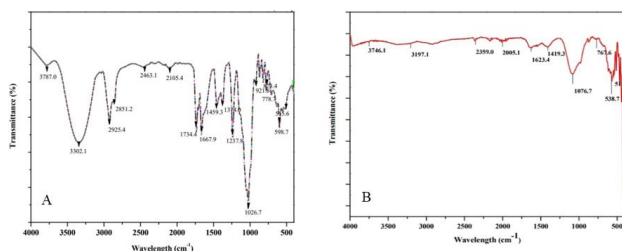


Figure 5. (a) FT-IR Spectrum of *Azadirachta indica* or neem fruit latex (NFL); (b) FT-IR Spectrum of NFL based FeO-MNPs.

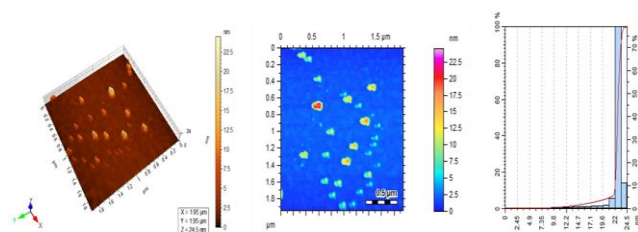


Figure 6. AFM images of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

3.5. Analysis through AFM Technique

AFM analysis showed that NFL based iron oxide NPS exhibited a size distribution between the ranges of 2.5 to 24.5 nm with average particles of 22.5 nm as depicts in **Figure (6)**.

3.6. Analysis through SEM and EDX

SEM analysis revealed that FeO-MNPs exhibited various shapes, predominantly irregular, with spherical and granular structure and cubic phase structure (close-packed structure), ranging in size from 20.59 to 84.58 nm with several agglomerates as shown in **Figure (7)** due to magnetic attraction generated by individual FeO-MNPs. Different plant-based iron oxide nanoparticles synthesis reported variations in size using SEM techniques like *Azadirachta indica* (neem) leaf extract has been utilized for one pot synthesis of iron oxide nanoparticles which exhibited irregular morphology and average size of 9–12 nm (Zambri et al., 2019^c). Black carrot (*Daucus carota* subsp. sativus) vegetable extracts were also employed which produced nanoscale particles that exhibited quasi-spherical shape, with varying sizes from 80-103 nm (Uncu Kirtiş et al., 2025). *Phyllanthus niruri* extract was utilized in another green approach synthesis spherical iron oxide nanoparticles with an average size of approximately 12.34 nm (VG, 2018). Similarly, *Hibiscus rosa-sinensis* flower extract was used in a microwave assisted synthesis method. The nanoparticles were in the range of 20-40 nm and mostly spherical in shape (Buarki et al., 2022). In our study we have introduced neem *Azadirachta indica* or neem fruit latex a novel green approach used for the iron oxide nanoparticles synthesis. EDX analysis was performed for further assessing the stoichiometry and element composition of the iron oxide nanoparticles as mentioned in **Table 2**. The appearance of clear peaks shown in **Figure (8)** at 6.4 and 7.2 keV associated with obligatory energies of iron (Fe) whereas, oxygen (O) peak appears strongly exhibits highest intensity at around 0.5 keV reinforcing the possibility of iron oxide structure in which Fe atoms are chemically bonded with O (Dhar et al., 2021 & Sirdeshpande et al., 2018).

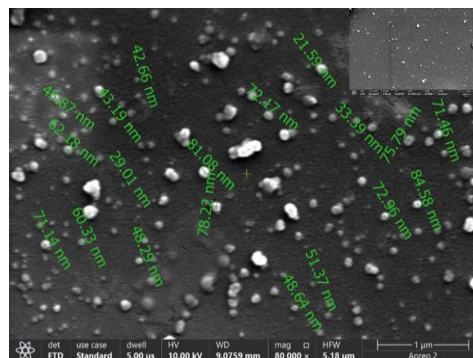


Figure 7. SEM results of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

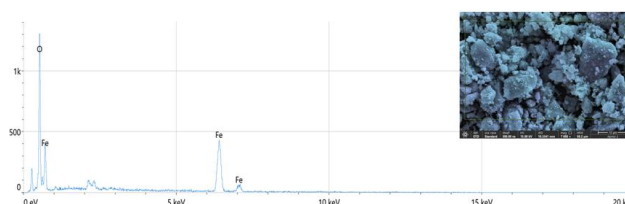


Figure 8. EDX Spectrum of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

Table 2. Elemental Composition of NFL based FeO-MNPs

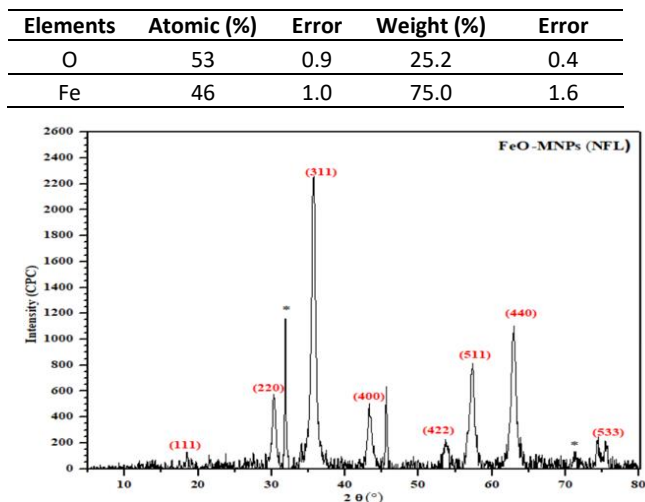


Figure 9. XRD Spectrum of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

3.7. Analysis by XRD

Figure 9 displays diffraction peaks at 18.6° , 30.4° , 35.7° , 43.3° , 53.6° , 57.3° and 63.0° corresponded to the (111), (220), (311), (400), (422), (511), and (440) planes respectively confirming the single-phase crystalline nature of the iron oxide nanoparticles matching exactly to the JCPDS No. 19-629. Usha *et al.*, 2022 & Majid *et al.*, 2021 obtained similar results in which distinct peak of Fe_3O_4 was observed at 30.47° (2θ) and intense peak at $2\theta = 35.96^\circ$. The XRD spectrum displayed the sharp and narrow diffraction peaks suggesting that the synthesized FeO-MNPs were both pure and highly crystalline. They exhibited a distinct rhombohedral crystalline nature, consistent with the findings reported by (Yazdani & Seddigh, 2016; Stan *et al.*, 2017 & Chinyerenwa *et al.*, 2018). The XRD pattern confirmed the crystalline size of synthesized FeO-MNPs and average particle size calculated by employing Debye-Scherrer formula was 16.5 nm.

3.8. Analysis by DLS

Dynamic Light scattering (DLS), measures the hydrodynamic size of nanoparticles which is generally larger as compared to microscopic size which represents only the actual core of nanoparticles. This increase in size is mainly due to particle aggregation in aqueous solution, which is a typical phenomenon for iron nanoparticles due to high surface energy and magnetic interactions which impacts their colloidal stability in water-based applications. Furthermore, the presence of biological material used as capping agents can also contribute to an increase in the hydrodynamic size. Polydispersity index (PDI) was found to be 2163 ± 0.455 d-nm with an intercept value of 0.930 whereas zeta potential was found -12.6 mV as shown in **Figure (10)**, which indicated moderate repulsive forces between the particles and reasonable stability in the colloidal state. Even though stability is typically greater when zeta potential is higher than $+20$ mV or less than -20 mV, however the measured value suggests a balance that prevents significant aggregation (Nwamezie & Igwe, 2018).

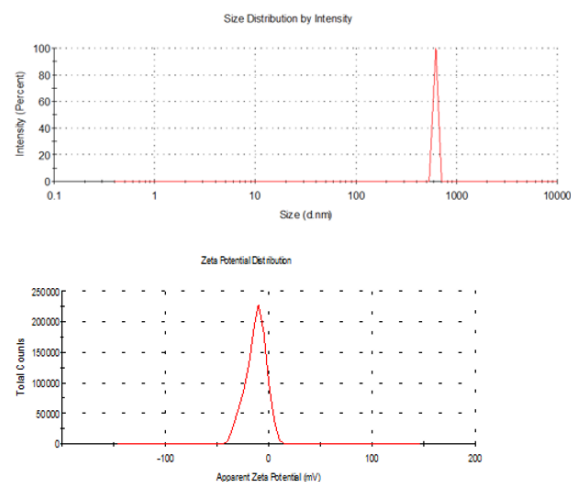


Figure 10. DLS Spectrum of *Azadirachta indica* or neem fruit latex (NFL) based FeO-MNPs

3.9. Adsorption Kinetic and Isotherms Modeling

The results showed that synthesized magnetic nanoparticles demonstrated favorable properties that positioned them to be used for treating wastewater contaminated with synthetic dyes. Following this, FeO-MNPs were further evaluated for their adsorption potential in removing reactive yellow 176 dye through adsorption kinetics and isotherm modeling. Reactive yellow 176 is an anionic dye containing negatively charged sulfonate groups which could interact with the surface hydroxyl groups of iron oxide nanoparticles particularly at lower pH promoting strong electrostatic attractions. The nanoparticles surface hydroxyl groups could also form hydrogen bonding with functional groups present in the dye, such as amide ($-\text{NH}_2$) and hydroxyl ($-\text{OH}$) groups. Also, the aromatic rings and conjugated azo ($-\text{N}=\text{N}-$) system in Reactive Yellow 176 could facilitate π - π stacking interactions with the iron oxide nanoparticle surface, further stabilizing dye adsorption.

Table 3. Effects of pH on removal efficiency & capacity of RY-176 dye with FeO-MNPs

pH	Co (mg L ⁻¹)	Removal %	Removal Capacity (mg g ⁻¹)
Acidic (4)	30	99.95	29.89
Neutral (7)	30	95.18	28.34
Basic (9)	30	19.27	5.84

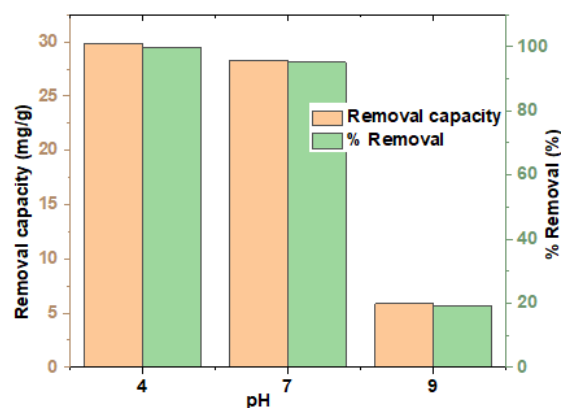


Figure 11. Effects of pH on removal efficiency & capacity of RY-176 dye with FeO-MNPs

3.10. Impact of Experimental Variables on the Adsorption Process

pH of Solution

The effect of pH on the removal efficiency of iron oxide nanoparticles for removal of RY-176 dye showed that acidic conditions enhanced both parameters, while alkaline conditions significantly hindered them. At an initial concentration of 30 mgL⁻¹, the percent removal was 99.95%, 95.18%, and 19.27% at pH 4, 7, and 9, respectively, with corresponding removal capacities of 29.89, 28.34, and 5.84 mg g⁻¹ as shown in **Table 3** and **Figure (11)**. The high removal capacity at pH 4 suggests favorable electrostatic interactions between the dye molecules at the nanoparticle surface. In contrast, the decrease at neutral pH and further decline at pH 9 indicated that an increasingly negative surface charge in alkaline conditions likely caused repulsion with the dye, reducing adsorption effectiveness. This trend highlights

the importance of acidic pH for optimizing iron oxide nanoparticles in dye removal applications.

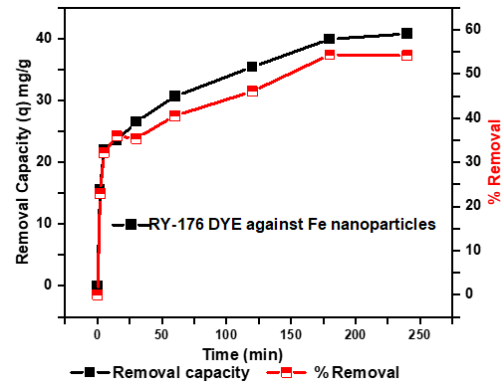


Figure 12. Effects of time on removal efficiency & capacity of RY-176 dye with FeO-MNPs

Table 4. Kinetic parameters of Pseudo-first-order (PFO) and Pseudo-second-order (PSO) models for the RY-176 dye removal

Material	Conc. mg L ⁻¹	Pseudo-first-order kinetic model			Pseudo-second-order kinetic model		
		q _e mmol g ⁻¹	k ₁ min ⁻¹	R ²	q _e mmol g ⁻¹	k ₂ mmol g ⁻¹ min ⁻¹	R ²
FeO-MNPs	30	33.340	0.212	0.82	36.499	0.006	0.90

Table 5. Removal of RY-176 dye using different adsorption models

Adsorption Models		FeO-MNPs
Langmuir		
Q _{max} (mmol g ⁻¹)		123.68±4.90
K _L (dm ³ mg ⁻¹)		0.027
R ²		0.986
Freundlich		
n		4.203±0.64
K _F [mmol g ⁻¹ (mg dm ⁻³) ^{-1/n}]		28.653±5.20
R ²		0.969
Sips		
Q _{max} (mg g ⁻¹)		128.793±17.15
K _s [(mg dm ⁻³) ^{-1/n}]		0.0385
n _s		1.127±0.36
R ²		0.987

Contact duration

The contact time between iron oxide nanoparticles and the dye is a critical factor in the dye removal process, impacting both the % removal and removal capacity. This was demonstrated in experiments using 30 mgL⁻¹ dye solution with 5 mg iron oxide nanoparticles at 298 K in acidic conditions across varying durations as shown in **Figure (12)**. In the initial 25 minutes, rapid removal was observed, with 35% of dye removed and a removal capacity of 25 mg/g, attributed to the abundance of active sites on the nanoparticles surface. Over time, the % removal and removal capacity gradually increased, achieving approximately 60% removal and 40 mg g⁻¹ removal capacity by 250 minutes. This slower removal rate in later stages is likely due to the reduction in available active sites and saturation effects, suggesting an optimal contact time of 250 minutes for maximum removal efficiency. Under optimal conditions at pH 4, the dye removal reached 99.95% with removal capacity of 29.89 mg/g, while increasing pH to 7 and 9 reduced the

removal efficiency to 95.18% and 19.27%, with corresponding removal capacities of 28.34 mg g⁻¹ and 5.84 mg g⁻¹, respectively.

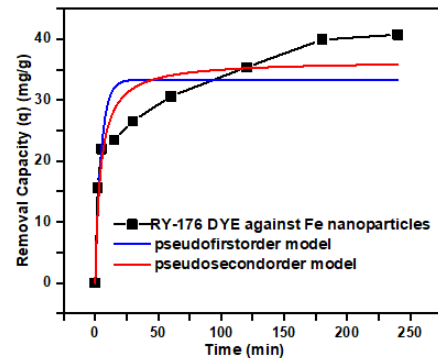


Figure 13. Kinetic isotherms for the removal of RY-176 dye and experimental data with adjustment to PFO and PSO Models

Adsorption kinetics

Kinetic studies offer valuable insights into the removal mechanism by modeling removal rates and fitting

experimental data into kinetic models. In this work, the removal kinetics of a 30 mg L⁻¹ dye solution with iron oxide nanoparticles at 298 K were analyzed using pseudo-first-order (PFO) and pseudo-second-order (PSO) models (Ahmed *et al.*, 2018; Agbovi & Wilson., 2021; Ahmed *et al.*, 2016). The equations for the kinetic models are provided below in equation 5 & 6.

$$(PFO) \ q_t = q_e (1 - e^{-k_1 t}) \quad (5)$$

$$(PSO) \ q_t = \frac{q_e^2 k_2 t}{1 + q_e k_2 t} \quad (6)$$

Where, q_e represents the removal capacity of the adsorbent (mg g⁻¹) at equilibrium, while q_t represents the capacity of adsorption at 't' time interval. The parameters k_1 and k_2 represent to the PFO and PSO rate constant, respectively. Plotting non- linear graph of q_t against t provides the values of the rate constant (k) and q_e as shown in **Table 4** and **Figure (13)**.

For the PFO model, the removal capacity at equilibrium (q_e) was found to be 33.34 mmol g⁻¹, with a rate constant K of 0.212 min⁻¹ and a correlation coefficient R^2 of 0.82, suggesting a moderate fit. Conversely, the PSO model provided a slightly higher (q_e) of 36.49 mmol g⁻¹ and a lower rate constant K of 0.006 mmol g⁻¹ min⁻¹, with an improved R^2 of 0.90. The higher correlation coefficient in the PSO model, suggests a more accurate depiction of the removal process. The better fit of the PSO model implies that the degradation rate is controlled by chemisorption, involving valence forces through sharing or exchange of electrons between dye molecules and active sites on the nanoparticles. This interaction leads to a slower, yet more stable, degradation process as equilibrium is reached. Therefore, the PSO model highlights the strong chemical affinity involved in the adsorption process and provides a more reliable framework for predicting dye removal efficiency over extended contact times in aqueous system.

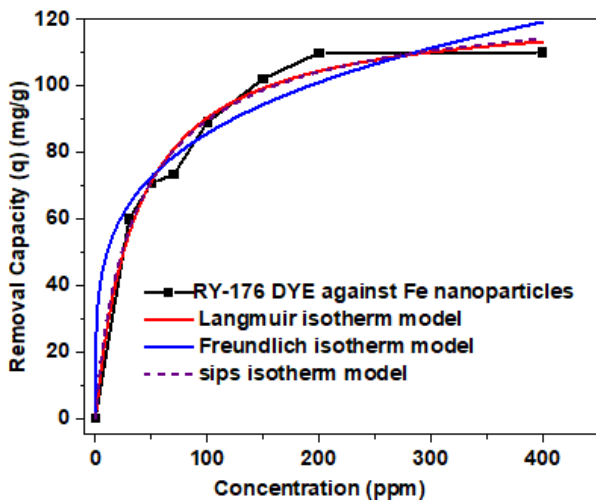


Figure 14. Adsorption Isotherms with adjustment to Langmuir, Freundlich and Sips Models

Adsorption Isotherm

The removal equilibrium investigates the distribution of dye molecules between iron oxide nanoparticles and the

solution when the equilibrium is attained. This analysis considers the concentration of the dye to understand the extent to which it partitions between the two phases. In this study, the Langmuir, Freundlich, and Sips isotherms, well-known models, were explored across a range of dye concentrations. The objective was to elucidate the maximum removal capacity (q_{max}) and the removal mechanism of the nanoparticle. The Langmuir adsorption isotherm, which assumes monolayer adsorption on a homogeneous surface, was employed for this purpose as mentioned in equation 7 (Sayed *et al.*, 2024; Ali *et al.*, 2024).

$$q_e = \frac{q_m K_L C_e}{1 + K_L C_e} \quad (7)$$

Here, q_m (mg g⁻¹) denotes the highest adsorption capacity, indicating the maximum quantity of dye that can be adsorbed per unit mass of adsorbent. K_L (dm³ mg⁻¹), the Langmuir constant, reflects the affinity of the nanoparticles for the dye. The Freundlich isotherm, on the other hand, describes multilayer adsorption occurring on a heterogeneous surface, as mentioned below in equation 8.

$$q_e = K_F C_e^{\frac{1}{n}} \quad (9)$$

In this equation, q_e denotes the removal equilibrium capacity (mg g⁻¹), while C_e represents the equilibrium concentration after removal. K_F is the Freundlich constant. The Sips adsorption model, which is frequently employed in adsorption studies, combines aspects of both the Langmuir and Freundlich models. It can be mathematically expressed as mentioned in equation 9.

$$q_e = \frac{Q_{max} K_s C_{eq}^{\frac{1}{n_s}}}{1 + K_s C_{eq}^{\frac{1}{n_s}}} \quad (10)$$

Where q_e denotes equilibrium removal capacity, measured in mg g⁻¹, while C_{eq} denotes the concentration of the dye at equilibrium in the solution, expressed in mg L⁻¹. K_s stands for the Sips constant and is measured in dm³ mg⁻¹.

The adsorption isotherms for dye removal by iron oxide nanoparticles indicate that each of the Langmuir, Freundlich, and Sips models provided useful insights into the removal process. In the Langmuir model, with a high $Q_{max} = 123.68$ mmol g⁻¹, $K_L = 0.27$, and $R^2 = 0.986$, the data suggest a strong affinity for dye molecules forming a monolayer on homogeneous surface sites, reflecting efficient initial removal adsorption that may support degradation as shown in **Table 5** and **Figure (14)**. The Freundlich model, with $n = 4.20$, $K_f = 28.65$, and $R^2 = 0.969$ further confirmed this process with a favorable adsorption intensity $n > 1$, indicating multilayer adsorption on heterogeneous surfaces, which might enhance interaction sites for degradation reactions. The Sips model, which is Langmuir and Freundlich models hybrid, yielded a slightly higher Q_{max} of 128.79 mg g⁻¹, with $K = s =$

0.038, $ns = 1.127$, and $R^2 = 0.987$, showing the system's adaptability to both low and high dye concentrations. The strong correlation values (close to 1) in each model suggest high reliability in describing the dye removal mechanism, with the Sips model slightly outperforming in fitting data. The findings suggest that the dye removal by iron oxide nanoparticles occurred through a two-step process: an initial adsorption phase followed by degradation mediated via surface interactions.

3.11. Multiple Reusability of FeO-MNPs

The reusability of adsorbent FeO-MNPs is important to decrease the production cost. The percentage of removal decrease from cycle 1 to 3 with following the conditions like pH 4, concentration 30 mg L⁻¹, dosage of adsorbate 5 mg and time 150 mins. The removal of RY176 dye achieved up to 95.45% in 1st cycle after 3rd run removal efficiency decreased to 78.52% as mentioned in **Figure (15)** due to blockage of the active sites of adsorbate. During recovery phase, it is found very easy to recover the FeO-MNPs with minimal loss due to its magnetic nature during recovery phase.

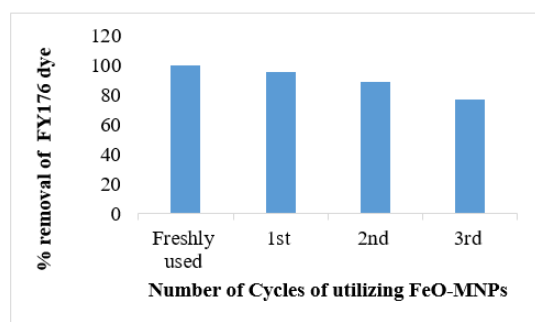


Figure 15. Multiple reusability of FeO-MNPs for RY 176 dye removal

4. Conclusion

This latex-based biosynthetic method for FeO-MNPs offers a sustainable, cost-effective route for dye remediation, with promising implications for industrial wastewater treatment. The nanoparticles demonstrated high reactivity in effectively removing reactive yellow 176 dye. The careful synthesis and application of these nanoparticles for scouring pollutants from industrial effluents underscore their promise as an eco-friendly solution for wastewater treatment. *Azadirachta indica* (Neem) fruit latex substantiates the presence of a variety of chemicals including phenols, alkaloids etc. mediates in reduction process and act as capping agent in the synthesis of FeO-MNPs. Furthermore, this study showed that Pseudo-second order (PSO) is the best fitted adsorption kinetics indicating chemisorption controlled process whereas the Sips model provided the best fitting adsorption isotherm as compared to Langmuir and Freundlich. Overall, FeO-MNPs found to be an alternative option for wastewater treatment

Highlights

- The synthesis of FeO-MNPs was carried out using a green approach without using harmful chemicals as reducing, capping or stabilizing agents.

- Latex of *Azadirachta indica* or unripe neem fruit was used in this green synthesis resulting in stable nanoscale particles between 2.5-24.5 nm in size by Atomic Force Microscopy (AFM).
- Synthesized nanoscale particles was examined for their effectiveness against Reactive Yellow RY-176 dye, revealing an impressive removal efficiency of 99 % at pH 4.
- The reusability efficiency was 78.52 % after three cycles.

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