

Preparation and application of silver oxide/sawdust nanocomposite for Chromium (VI) ion removal from aqueous solutions using column system

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Received: 20/12/2016, Accepted: 20/08/2017, Available online: 18/10/2017

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Abstract

This study describes the removal of Cr (VI) ions from aqueous solutions using silver oxide/sawdust nanocomposite (Ag₂O/SD NC) that was prepared using the chemical precipitation method. Scanning Electron Microscopy (SEM), X-ray Diffraction (XRD) and Fourier Transform Infrared Spectroscopy (FT-IR) techniques were used for characterization of Ag₂O/SD NC. The effect of various parameters such as feed solution pH (2, 6, and 10), influent concentration of Chromium (VI) (25-100 mg/L), flow rate (1-5 mL/min) and bed depth (4-12 cm) on breakthrough curves (BTC) were investigated. Three well-known column adsorption kinetics models including Thomas, Adams-Bohart and Bed Depth Service Time (BDST) models were applied to fitting the column experimental data. Based on Thomas model, the equilibrium adsorption reached 13.41 mg/g when a Cr (VI) polluted solution with influent concentration of 25 mg/L Cr (VI) passed through the column with a flow rate of 2 mL/min. Desorption studies reveal that recovery of uploaded Cr (VI) from the Ag₂O/SD NC adsorbent or exhausted column regeneration can readily be achieved using a dilute solution of NaOH (0.01 M) as eluent. This study indicated that the Ag₂O/SD NC can be used as an effective, efficient and environmentally friendly adsorbent for the removal of Cr (VI) ions from water resources.

Keywords: Removal, Cr (VI), Silver oxide/sawdust Nanocomposite, Column system

1. Introduction

The presence of toxic heavy metals in water creates serious health and environmental hazards. Hence, significant research activities have been aimed to remove heavy metal ions from water resources. Chromium stands as one of the most toxic and carcinogenic chemical element, has been commonly used in a number of industrial processes, such as leather tanning, cement industries, production of steel and other metal alloys, electroplating, metal polishing, photographic material and corrosive paints (Badruddoza *et al.*, 2013; Li *et al.*, 2006; Malkoc *et al.*, 2006). Cr (VI) is a

powerful carcinogenic agent that modifies the DNA transcription process causing important chromosomal aberrations. Due to high toxicity and bioaccumulation of Cr (VI), the removal of Cr (VI) ions from contaminated aqueous solution is the great importance. Based on the World Health Organization (WHO), maximum permissible concentration of Cr (VI) ions in drinking water has been reported 0.01 mg/L, whereas US Environmental Protection Agency (EPA) has set the maximum contaminate level for Cr (VI) at 0.1 mg/L (Lv *et al.*, 2011; Wang *et al.*, 2013).

Various methods for removing of Cr (VI) have been developed, such as cyanide treatment, electro-chemical precipitation (Liu *et al.*, 2011), reverse osmosis, adsorption (Bayramoglu and Arica, 2011), and ion exchange (Sardohan *et al.*, 2010). Among these methods, adsorption is the most widely used method because of the simplicity of operation, high efficiency and low cost (Li *et al.*, 2013; Rangabhashiyam *et al.*, 2016a; Saranya *et al.*, 2017). The adsorbent materials must have a high specific surface area, many adsorption sites, and chemical stability. Batch reactors were easy to use in the laboratory study, but less convenient for industrial applications. On the other hand, fixed-bed columns were widely used in various chemical industries for their operation. The performance of packed beds is described through the concept of the breakthrough curve. The time for breakthrough appearance and the shape of the breakthrough curve are very important characteristics for determining the operation and the dynamic response of an adsorption column (Chen and Wang, 2000).

In the present study, Ag₂O/SD NC was employed as a low-cost and efficient nano adsorbent material for removal of Cr (VI) ions from aqueous solutions. Chemical modification can improve the sorption capability of wood sawdust for removal of any toxic materials from aqueous solutions (Moafi *et al.*, 2016). Continuous adsorption experiments were performed to see the effect of influencing parameters such as feed solution pH, flow rate, initial concentration of Cr (VI) solution and bed height of adsorbent on Cr (VI) removal. Mathematical models for fixed-bed adsorption systems, including Thomas, Adams-Bohart and Bed depth

service time (BDST) have been employed for determination of the column performance. Analysis of Cr (VI) was carried out spectrophotometrically using 1, 5-diphenyl-carbazide (DPC) reagent (Carrington *et al.*, 2007; Gajghate *et al.*, 1992). DPC reagent dissolved in acetone solvent and sulfuric acid (pH=2) and the intense adsorption of the red-violet colored complex formed ($\lambda_{\text{max}}=540$ nm). A calibration curve based on Beer Lambert law was employed for the analysis of unadsorbed Cr (VI) in supernatants.

2. Experimental

2.1. Materials and Methods

All chemicals used were of analytical grade (AR) and were prepared in distilled water. Sawdust samples (SD) from dicer were obtained from a local carpentry workshop. Silver nitrate (AgNO_3), sodium hydroxide (NaOH), Potassium dichromate ($\text{K}_2\text{Cr}_2\text{O}_7$), 1,5-di-phenylcarbazide reagent was obtained from Merck. A column made of Pyrex glass tube having an inner diameter of 1.0 cm and height of 40 cm was used for adsorption experiments in the fixed bed column system. A Metrohm pH meter (Model 827) with a combined double junction glass electrode, calibrated against two standard buffer solutions at pH 4.0 and 7.0, was used for pH measurements. The pH adjustments of the test solutions were made using either dilute NaOH or HCl solutions. A single beam Perkin-Elmer UV-Vis spectrophotometer with a 1 cm cell was used for measuring all of the adsorption data. Philips PW X-ray diffractometer at a voltage of 40 kV and a current of 100 mA coupled with monochromatic $\text{Cu-K}\alpha$ radiation ($\lambda=1.5405$ Å), with a scan rate of $2^\circ/\text{min}$ for 2θ between 10 and 70 degrees. Scanning electron microscopy (SEM) (LEO 1430VP) was used to characterize the surface morphology of the adsorbent. The infrared spectra ($400\text{--}4000$ cm^{-1}) of the samples were recorded on an alpha FT-IR (2011 Bruker Optic GmbH) instrument.

2.2. Preparation of adsorbent

5.0 g of sawdust was added to 50 mL of silver nitrate (0.20 M) solution and then 50 mL of sodium hydroxide (0.40 M) was added slowly and drop wise, while the solution was constantly stirred with a magnetic stirrer, until the solution had the consistency of a brown-black colloidal suspension. After addition of NaOH solution, the mixture was heated for 2 hours at temperature about $70\text{--}80$ °C accompanied by stirring using a magnetic stirrer. Then, the solution allowed cooling gradually. The coated particles separated from the solution by filtering, washed with distilled water and dried in oven at $90\text{--}100$ °C until constant weight.

2.3. Adsorption studies

Continuous flow adsorption studies were conducted in a column made of Pyrex glass tube. Before pouring the adsorbent into column, a layer of glass wool was placed to prevent loss of adsorbent. A stock Cr (VI) solution (200 mg/L) was prepared in distilled water using $\text{K}_2\text{Cr}_2\text{O}_7$. All working solutions were prepared by diluting the stock solution with distilled water. The feed solution was poured from above at a desired flow rate. To determine the influences of initial chromium (VI) concentration, flow rate

and bed height, different experiments were performed. The plan of the experiments is given in Table 1.

The breakpoint time is defined as the time of adsorption when the outlet Chromium (VI) concentration from the column was about 1-5% of the inlet concentration.

Table 1. Experimental plan for adsorption of Cr (VI) onto $\text{Ag}_2\text{O}/\text{SD}$ NC

Parameters		
Initial Cr (VI) concentration (mg/L)	Bed height (cm)	Flow rate (mL/min)
25	4	2
50	4	2
100	4	2
25	8	2
25	12	2
25	4	1
25	4	5

The total stoichiometric quantity of adsorbed Cr (VI) (q_{total}) in the column for a given inlet Cr (VI) concentration and flow rate is calculated using Eq. 1 (Rangabhashiyam *et al.*, 2016b).

$$q_{\text{total}} = \frac{FA_c}{1000} = \frac{F}{1000} \int_{t=0}^{t=t_{\text{tot}}} C_{\text{ad}} dt \quad (1)$$

The area under the breakthrough curve (A_c) is obtained by plotting the adsorbed concentration (C_{ad}) versus time (t). The total amount of Cr (VI) sent to column (m_t) was calculated using Eq. 2.

$$m_t = \frac{C_{\text{ad}} F t_t}{1000} \quad (2)$$

Where, t_t is the time equivalent to the total stoichiometric capacity of the column which is defined by Eq. 3.

$$t_t = \frac{q_t}{m_t} t_f \quad (3)$$

The removal percentage (S) was calculated using Eq. 4 (Gupta and Babu, 2010).

$$S(\%) = \frac{q_t}{m_t} \times 100 \quad (4)$$

The regeneration efficiency (RE %) was calculated using Eq. 5.

$$\%RE = \frac{q_{\text{reg}}}{q_{\text{org}}} \times 100 \quad (5)$$

Where, q_{reg} is the adsorptive capacity of the regenerated column and q_{org} is the sorption capacity of the original column (Ansari *et al.*, 2013).

2.4. Mathematical modeling of breakthrough curves (BTC)

2.4.1. Thomas kinetics model

The data obtained in column in continuous mode studies was used to calculate the maximum concentration of Cr (VI) on adsorbent and the adsorption rate constant using the kinetic model developed by Thomas. The Thomas solution is one of the most general methods that is widely used for column performance determination. The expression by Thomas for an adsorption column is given as follows (Karimi *et al.*, 2012; Han *et al.*, 2006).

$$\frac{C_t}{C_0} = \frac{1}{1 + \exp\left[\frac{K_{Th}}{F}(q_0 \cdot x - C_0 \cdot V_{eff})\right]} \quad (6)$$

The linearization of Eq. 6 yields:

$$\ln\left(\frac{C_0}{C_t} - 1\right) = \frac{K_{Th} \cdot q_0 \cdot x}{F} - K_{Th} \cdot C_0 \cdot t \quad (7)$$

From the linear dependence of $\ln\left(\frac{C_0}{C_t} - 1\right)$ versus V , the adsorption removal capacity (q_0) and rate constant (K_{Th}) can be determined (Hasanzadeh *et al.*, 2016; Rangabhashiyam *et al.*, 2016c).

2.4.2. Bohart–Adams model

The Adams–Bohart model is based on the surface reaction theory that established a fundamental equation for describing the relationship between C_t/C_0 and t . The adsorption rate is controlled by external mass transfer and equilibrium is not instantaneous (Bohart and Adams, 1920). Generally, it is descriptive to analyze the initial part of a breakthrough curve ($C_t/C_0 < 0.5$) (Bhaumik *et al.*, 2013; Wang *et al.*, 2015). The expression is as follows:

$$\ln\left(\frac{C_t}{C_0}\right) = K_{AB} C_0 t - K_{AB} N_0 \left(\frac{Z}{U_0}\right) \quad (8)$$

Where U_0 is the superficial velocity (cm/min) defined as the ratio of the volumetric flow rate q (cm³/min) to the cross-sectional area of the A bed (cm²).

2.4.3. Bed-depth service time model (BDST)

It is generally accepted that the BDST model offers the simplest approach and rapid prediction of adsorbent design and column performance. The BDST is a model for predicting the relationship between bed depths (z) and service time (t). This BDST model was focused on the estimation of characteristic parameters such as the maximum adsorption capacity (N_0) and kinetic constant (K_a). This model assumes that the adsorption rate is proportional to the residual capacity of the sorbent and the concentration of the sorbing species. The service time was related to process conditions and operating parameters as (Negrea *et al.*, 2011):

$$\ln\left(\frac{C_0}{C_b} - 1\right) = \ln\left(\exp\left(\frac{K_a \cdot Z \cdot N_0}{F}\right) - 1\right) - K_a C_0 t_B \quad (9)$$

A linear relationship between bed-depth and service time may be given by Eq. 10 (Rangabhashiyam *et al.*, 2016 b).

$$t_B = \frac{N_0}{C_0 U_0} Z - \frac{1}{K_a C_0} \ln\left(\frac{C_0}{C_b} - 1\right) \quad (10)$$

Thus, N_0 and K_a can be evaluated from the slope ($\frac{N_0}{C_0 U_0}$)

and the intercept ($-\frac{1}{K_a C_0} \ln\left(\frac{C_0}{C_b} - 1\right)$) of the plot of t_B versus Z , respectively.

3. Results and discussion

3.1. Characterization of the adsorbent

3.1.1. X-ray Diffraction (XRD) Analysis

The XRD patterns of Ag₂O nanoparticles, untreated Sawdust and Ag₂O/Sawdust NC before Cr (VI) adsorption prepared via a chemical precipitation method, are shown in Figure 1. The XRD pattern of Ag₂O NPs and sawdust substrate are presented in the Figure 1a, 1b, respectively. In Figure 1a, the bulk of the X-ray signal originated just from the untreated sawdust substrate in $2\theta = 23^\circ$. The XRD pattern for the Ag₂O nanoparticles (Figure 1b) had shown reflections and planes at $2\theta = 32.94^\circ$ (111), 38.22° (200), 55.12° (311) and 65.70° (222) with cubic system and confirms that Ag₂O NPs was synthesized properly. The characteristic peaks ascertained from the XRD pattern of untreated Ag₂O/SD NC (Figure 1c) indicates that both silver and silver oxide nanoparticles are presented in the prepared nanocomposite (Sullivan *et al.*, 2013; Moafi *et al.*, 2016). It seems that some silver nanoparticles have also been produced in the course of synthesis of Ag₂O/SD NC due to the presence of some reductive functional groups in the sawdust. The reflections observed at $2\theta = 37.90^\circ$ (111), 43.90° (200) and 64.18° (220) are related to the face-centered cubic (FCC) structure of silver oxide and metallic silver particle with the (111) reflection intensified considerably, in the presence of sawdust (Yang *et al.*, 2004; Janardhanan *et al.*, 2009).

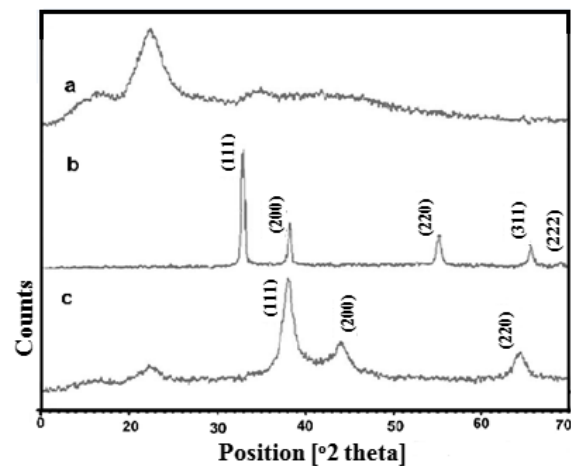


Figure 1. XRD patterns of a) Untreated Sawdust, b) Ag₂O nanoparticles and c) Ag₂O-coated sawdust

3.1.2. Scanning Electron Microscopy (SEM) Analysis

SEM analysis was performed for characterizing of the surface morphology and particle sizes of the synthesized adsorbent ($\text{Ag}_2\text{O}/\text{SD NC}$). SEM micrographs of the samples (untreated Sawdust, $\text{Ag}_2\text{O}/\text{SD NC}$ before and after Cr (VI) adsorption) are represented in Figure 2. It is evident that sawdust is a smooth and layered material without any pores (Figures 2a). In Figure 2b, the surface morphology of the chemically modified sawdust with silver oxide or silver NPs before Cr (VI) adsorption possessing a large amount of tiny particles with various sizes (53–65 nm) has been represented. An internal surface and pores are necessary

for efficient biosorption (Rangabhashiyam *et al.*, 2014). Figure 2c shows the SEM image of modified sawdust with Ag_2O NPs after adsorption of Cr (VI). SEM images were shown that the most of Ag_2O particles are irregularly shaped and relatively spherical with dimensions less than 160 nm. It can clearly be seen that the surfaces of adsorbent were covered with the fine particles of Cr (VI) and the bulk particles are separated. Moreover, the new nano-sized formed particles on the surface of the modified adsorbent ($\text{Ag}_2\text{O}/\text{SD}$) explains the high specific pore surface area and the enhanced sorption capacity of the modified adsorbent (Moafi *et al.*, 2016).

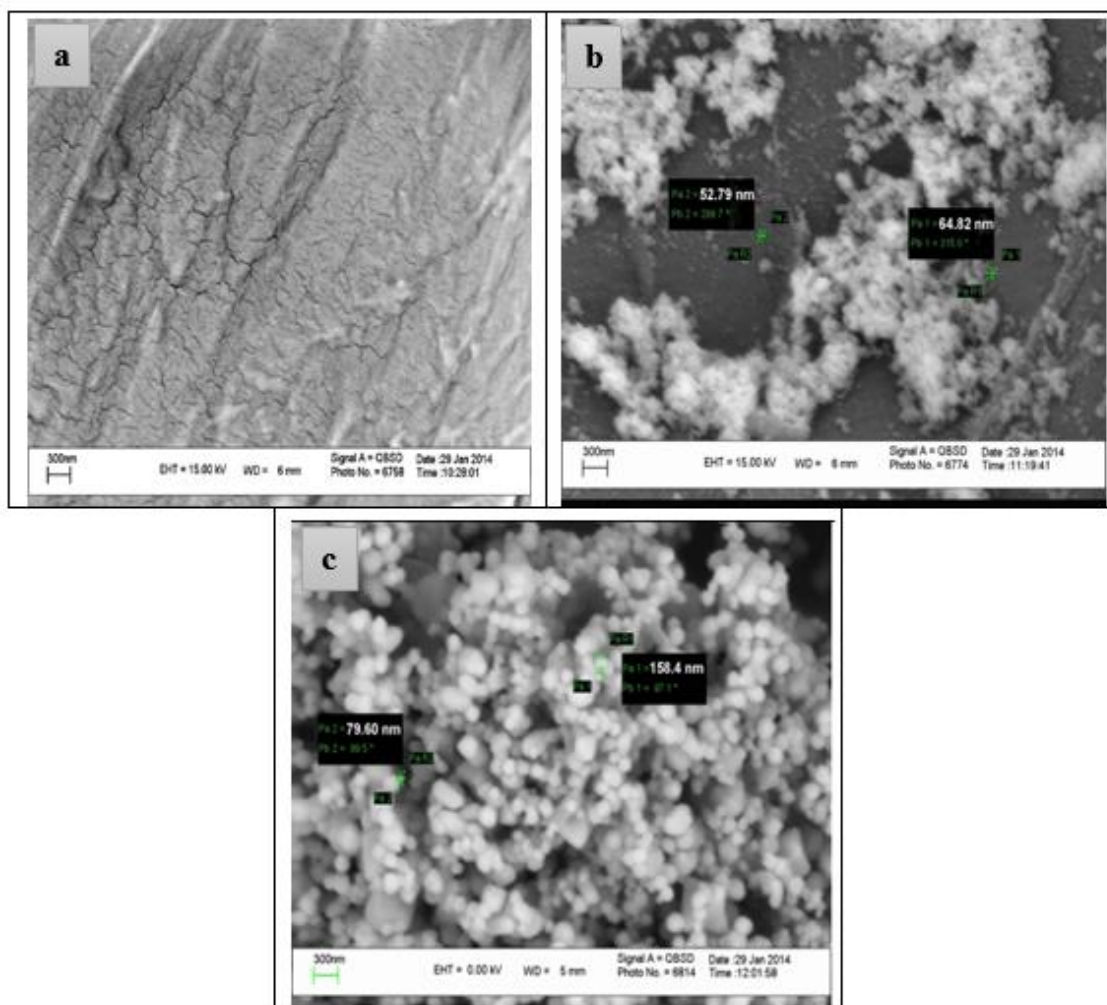


Figure 2. SEM micrograph of a) Untreated Sawdust b and c) $\text{Ag}_2\text{O}/\text{Sawdust}$ Nanocomposite before and after Cr (VI) Adsorption, respectively

3.1.3. Fourier Transform Infrared Spectroscopy (FT-IR) spectroscopy study

The FTIR spectra of $\text{Ag}_2\text{O}/\text{SD NC}$ before and after adsorption are shown in Figure 3. FTIR spectrum analysis confirms that the silver oxide nanoparticles are dispersed onto the sawdust. Figure 3a shows the FT-IR spectra of $\text{Ag}_2\text{O}/\text{SD NC}$. Sawdust bed in nanocomposite has a fiber structure and its main component is cellulose, which has a straight chain structure and large molecular mass. The adsorption peak at 3417 cm^{-1} indicates the hydroxyl groups and the contribution from C-H stretching, respectively

(Nakkeeran *et al.*, 2016a). The band at 2354 cm^{-1} is assigned to stretching vibrations of N-H or C=O groups probably due to amines and ketones, 1642 cm^{-1} is an indication of COO, C=O, and can also indicate the bending vibration of adsorbed water (Rangabhashiyam., 2015a), 1030 cm^{-1} is stretching vibration of C-O-C and O-H of polysaccharides in sawdust (Gode *et al.*, 2008). The absorption peak at 1416 cm^{-1} could be assigned to CO_2 stretching vibration (carbon dioxide from the atmosphere, perhaps due to the mesoporous nature of Ag_2O nanomaterials). The Ag-O-Ag stretching vibration band could be appeared at around of 607 cm^{-1} which is covered

due to the overspread of spectra in the region 400-500 cm^{-1} (Rahman *et al.*, 2013). FT-IR spectra of $\text{Ag}_2\text{O}/\text{SD}$ NC after adsorption of Cr (VI) are shown in Figure 4b. Two new peaks were observed in the FTIR spectra of Cr (VI) - loaded sorbents (807 and 900 cm^{-1}) attributed to Cr=O and Cr-O bonds of HCrO_4^- , which suggests that Cr (VI) was adsorbed on the surface of $\text{Ag}_2\text{O}/\text{SD}$ nanocomposite (Deng and Bai, 2004; Kyzas *et al.*, 2009). The uptake of Cr(VI) ions using some bioadsorbents has also been well evidenced employing EDX technique (Nakkeeran *et al.*, 2016b; Rangabhashiyam and Selvaraju., 2015b)

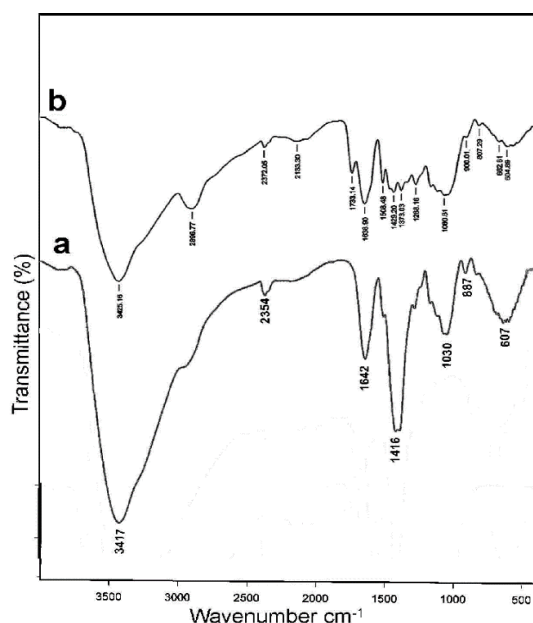


Figure 3. FT-IR spectra of a) $\text{Ag}_2\text{O}/\text{SD}$ NC before Cr (VI) adsorption, b) $\text{Ag}_2\text{O}/\text{SD}$ NC after Cr (VI) adsorption.

3.2. Effect of Operational Parameters

3.2.1. The effect of initial solution pH

The pH of feed solution is an important controlling parameter in the Cr (VI) adsorption process and thus the role of hydrogen ion concentration was examined from solutions at different pH, covering about 2.0, 6.0 and 10.0. Figure 4 shows the effect of pH values on adsorption of Cr (VI) ions onto $\text{Ag}_2\text{O}/\text{Sawdust}$ NC using a plot of dimensionless concentration (C_t/C_0) versus effluent volume (V). The initial pH of the adsorption medium relates to the adsorption mechanisms onto the adsorbent surface from water and reflects the nature of the physicochemical interaction of the species in solution and the adsorptive sites of adsorbent (Gao *et al.*, 2009). There are several forms of Cr (VI) ions existing in the solution; they are chromate (CrO_4^{2-}), dichromate ($\text{Cr}_2\text{O}_7^{2-}$) and hydrogen chromate (HCrO_4^-). These ions forms are related to the solution pH and total Cr (VI) concentration (Weckhuysen and Wachs, 1996; Rangabhashiyam and Selvaraju, 2015c). At pH lower than 6.8, HCrO_4^- is the conquering species, and at pH above 6.8 only CrO_4^{2-} is the most prominent. When the pH is low, they could be adsorbed by positively charged through electrostatic attraction (Li *et al.*, 2013; Rangabhashiyam, 2016d).

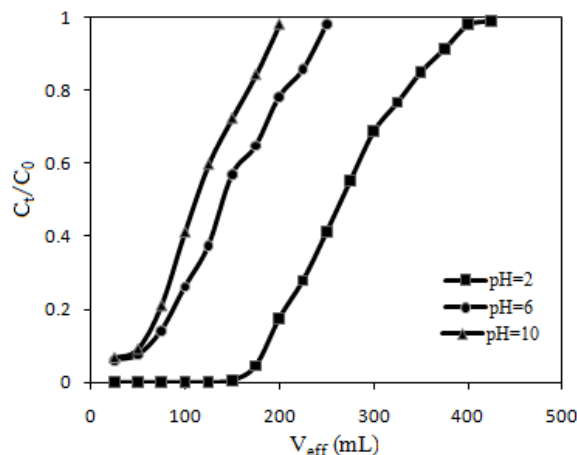


Figure 4. Effect of solution pH on BTC ($C_0 = 25 \text{ mg L}^{-1}$, F.r.= 2 mL min^{-1} , bed height = 4.0 cm, weight of adsorbent = 0.50 g)

As shown, the adsorption of Cr (VI) ions onto $\text{Ag}_2\text{O}/\text{SD}$ NC greatly affects the pH. The highest bed capacity and the longest breakthrough time were obtained at low pH value (pH=2).

3.2.2. The effect of flow rate

The effect of flow rate on the performance of the column was studied at three different flow rates of 1, 2 and 5 mL/min . During the experiments the bed height and initial concentration of feed was maintained at constant values of 4 cm and 25 mg/L , respectively. The breakthrough curves are depicted in Figure 5. The effects of flow rate on adsorption results can be illustrated according to mass transfer concepts.

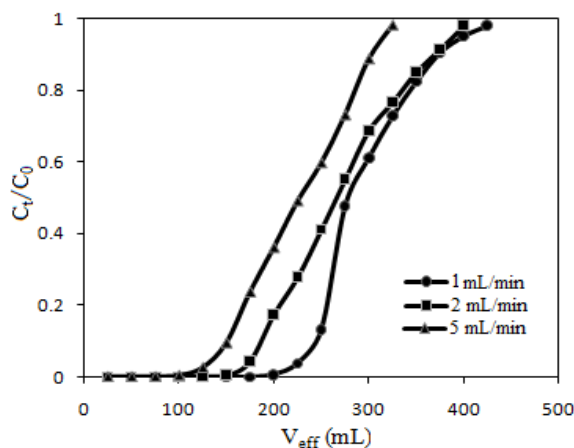


Figure 5. Effect of Flow rate on BTC ($C_0 = 25 \text{ mg/L}$, pH = 2.0, bed height = 4 cm (0.50 g))

As it could be expected with increasing of flow rate the contact time of the adsorbent and adsorbate is decreased. Therefore, the breakpoint has occurred sooner at higher flow rate since with increasing the flow rate, the saturation is happening faster too. When the flow rate is increased, the Chromium (VI) ions do not have enough time to penetrate and diffuse deeply into the pores. Therefore, equilibrium does not take place, and lower adsorption capacity might be achieved at higher flow rate of inlet

(Ansari *et al.*, 2013; Karimi *et al.*, 2012). As the data show (Figure 5), the breakthrough time decreased from 125 min to 75 min when the flow rate changed from 1 to 5 mL/min. The adsorption capacity was also decreased from 13.6 to 10.7 mg/g. However, as indicated the same sorption capacity was observed for both 2 and 5 mL/min flow rates. Then Flow rate of 2 mL/min was selected as the optimum flow rate for further studies.

3.2.3. Effect of influent Cr (VI) concentration

To evaluate the effect of Cr (VI) concentration, the inlet concentration of Cr (VI) varied from 25 to 100 mg/L under the same operating conditions of bed depth (4 cm), pH value about 2 and flow rate of 2 mL/min. Figure 6 depicts three BTC of Cr (VI) removal using Ag₂O/SD NC at three different initial Cr (VI) concentrations.

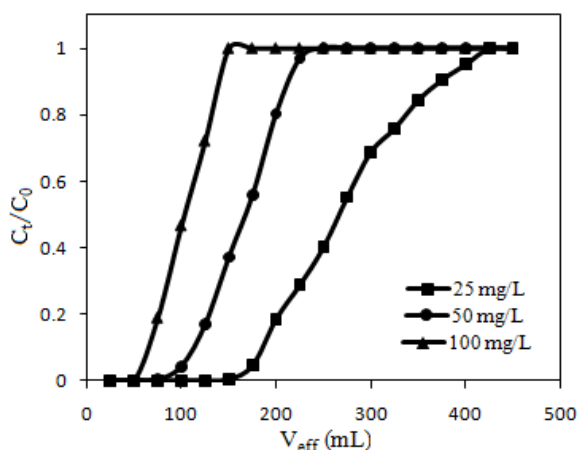


Figure 6. The effect of initial Cr (VI) concentration on BTC (pH=2.0, bed height =4 cm, F.r.=2 mL min⁻¹)

According to the results of Figure 6, the adsorbent gets saturated faster at higher initial concentration and breakthrough time (corresponding to 5% of influent concentration or $C_t/C_0 = 0.05$) decreases from about 100 min to about 35.5 min as the concentration of chromium (VI) solution increases from 25 to 100 mg/L. As shown (Fig. 6), more volumes of Cr (VI) polluted solution can be cleaned at lower concentrations, although some gradual increase was observed in the adsorption capacity of columns increased with increase in inlet initial concentration. Maximum bed capacities at 25, 50 and 100 mg/L of Cr (VI) concentrations were 12.9, 15.2 and 15.7 mg/g respectively. At lower inlet Cr (VI) concentration, BTC was dispersed and breakthrough occurred slowly. As inlet concentration increased, sharper BTC was obtained. These results demonstrate that the change of concentration gradient affects the saturation rate and breakthrough time.

3.2.4. Effect of column bed depth

The other factor that has the greatest effect on metal uptake capacity in a packed bed column is the height of adsorbent inside the column. To produce different bed heights, 0.50, 1.0, and 1.50 g of the Ag₂O/SD NC were added to yield 4, 8, and 12 cm height, respectively. During the experiments the flow rate and solution concentration

were maintained at constant values of 2 mL/min and 50 mg/L, respectively and the solution pH was kept at a value of about 2. The obtained breakthrough curves for the different bed heights are shown in Figure 7. As can be seen, the breakthrough point occurred more slowly when the bed height changed from 4 to 12 cm. The metal uptake capacity of the Ag₂O/SD NC and metal removal percentage increased with increasing the bed height. On the other hand, more volume of effluent can be treated with longer bed depth. However, the ultimate adsorption capacity of columns increased slightly (from 12.9 to 13.5 mg/g) when the bed height changed from 4 to 12 cm (Zhu *et al.*, 2012; Suganya *et al.*, 2016).

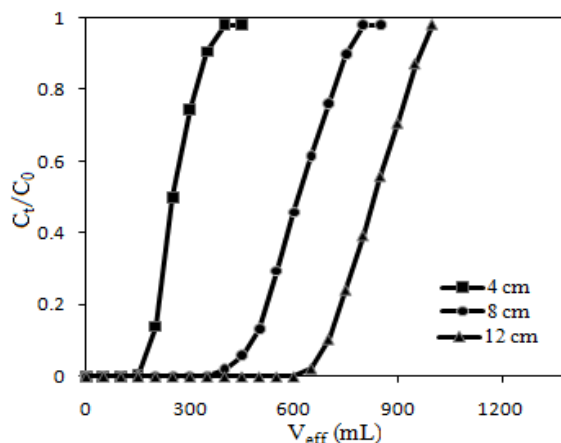


Figure 7. The effect of sorbent bed height on BTC ($C_0 = 25$ mg/L, pH = 2.0, F. r. = 2 mL min⁻¹)

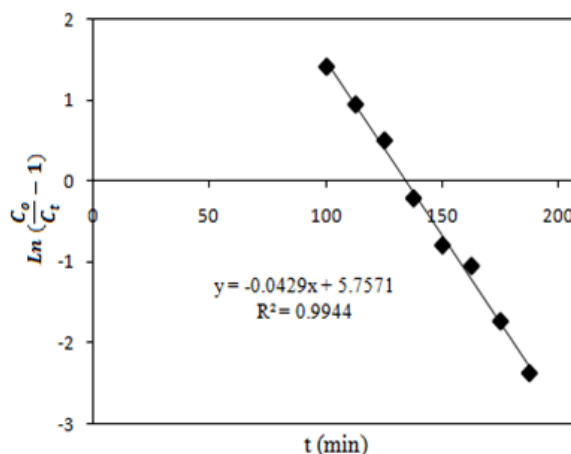


Figure 8. Comparison of the experimental BTC obtained at different initial concentrations of Cr (VI) according to the Thomas model (pH = 2.0, F.r.= 2 mL min⁻¹, bed height=4cm)

3.3. Breakthrough curve modeling

3.3.1. The Thomas model

Thomas model was applied to the experimental data with respect to influent concentration of Cr (VI), bed mass and flow rate conditions (Figure 8). A linear regression analysis was used in each set of data to determine the Thomas model parameters of q_0 and K_{Th} . The determined

coefficients were also obtained using linear regression analysis according Eq. 7. Based on the regression analysis ($R^2 = 0.9944$), the sorption data fit better with the Thomas kinetics model and it can be seen that with the bed mass increasing, the K_{Th} values decreased (Rangabhashiyam

et al., 2016c). The correlation coefficients were between 0.976 and 0.994 at the tested initial concentrations. The parameters calculated from the linear plot of Thomas model are summarized in Table 2.

Table 2. Thomas and Bohart–Adams model parameters obtained for Cr (VI) adsorption onto Ag₂O/SD NC under different experimental conditions.

Parameters	Thomas model				Bohart-Adams model		
	$K_{Th} \times 10^{-4}$ (mL/ mg. min)	q_0 (mg/g)	$q_{c(exp)}$ (mg/g)	R^2	$K_{AB} \times 10^{-4}$ (L/mg.min)	$N_0 \times 10^{-2}$ (mg/L)	R^2
Bed mass (g)							
0.5	1.72	13.41	12.90	0.9965	2.68	26.91	0.9791
1	1.42	15.29	13.50	0.9714	2.68	26.04	0.9880
1.5	1.46	13.73	13.80	0.9592	2.96	22.48	0.9880
Initial conc. (mg/L)							
25	1.72	13.41	12.90	0.9965	2.68	26.91	0.9791
50	1.65	15.62	15.20	0.9917	4.04	29.76	0.9705
100	1.63	16.31	15.70	0.9914	4.73	29.77	0.9670
Flow rate (mL/min)							
1	2.37	7.34	13.60	0.9713	3.00	13.59	0.9191
2	1.72	13.41	12.90	0.9965	2.68	26.91	0.9791
5	2.45	28.01	10.70	0.9481	6.04	52.59	0.9727

Table 3. Comparison of the column equilibrium adsorption capacity of different adsorbents with Ag₂O/SD NC nanocomposite.

Adsorbent	pH	Capacity (mg/g)	References
Ficus auriculata leaves powder	2.0	13.33	(Rangabhashiyam <i>et al.</i> , 2015c)
Strychnos nux-vomica shell	2.0	90.9	(Nakkeeran <i>et al.</i> , 2016)
poly 3-methyl thiophene / Sawdust	1.0	0.03	(Ansari and Fallah Delavar, 2010)
PPy/Sawdust	5.0	3.4	(Ansari and Khoshbakht Fahim, 2007)
Sawdust	<3	1.0	(Ansari and Khoshbakht Fahim, 2007)
Titanium (IV) oxide nanoparticle	2.0	12.94	(Debnath <i>et al.</i> , 2010)
PPy/Fe ₃ O ₄ nanocomposite	2.0	230.17	(Bhaumik <i>et al.</i> , 2013)
Rice Husk Carbon	2.0	38.1	(Singh and Singh, 2012)
Cetyl tri methyl ammonium	1.15-1.39	0.366	(Rajesh <i>et al.</i> , 2007)
Carbon slurry	2.08	28	(Gupta <i>et al.</i> , 2010)
Jute fiber (PANI-jute)	3.0	4.66	(Kumar and Chakraborty, 2009)
Olive oil industry waste	1-2	3.33	(Malkoc <i>et al.</i> , 2006)
Maize corn cob	2.0	0.28	(Garg <i>et al.</i> , 2007)
Rice straw	2.0	3.15	(Gao <i>et al.</i> , 2008)
Coal	2.1	6.73	(Gupta <i>et al.</i> , 1997)
Red mud	2.0	4.36	(Gupta <i>et al.</i> , 2001)
Cactus leaves	2.0	7.08	(Dakiky <i>et al.</i> , 2002)
Jatropha oil cake	2.0	0.82	(Garg <i>et al.</i> , 2007)
pine leaves	5.0	0.470	(Aliabadi <i>et al.</i> , 2006)
Eichhornia crassipes	2.0	5.6	(Saraswat and Rai., 2010)
Ag ₂ O/Sawdust nanocomposite	2.0	13.41	Current study

As shown (Table 2), with increasing the inlet of Cr (VI) concentration, the value of q_0 increased and the value of K_{Th} is decreased. In addition, although the deviations of experimental and predicted values were quite high, there was a good agreement between the experimental and predicted bed capacities.

To assess the feasibility of the Ag₂O/SD nanocomposite, comparison of the adsorption capacity of different low cost and commercially available adsorbents is shown in Table 3. When compared with other adsorbents (Table 3), the

results obtained from the breakthrough curve analysis of the present study indicate that Ag₂O/sawdust nanocomposite showed higher sorption capacity toward Cr (VI) and can be used for removal of Cr (VI) from aqueous solution more effectively.

3.3.2. Bohart-Adams model

Figures 5–7 show the fitting breakthrough curves with respect to bed mass, initial Cr (VI) concentration and flow rate by the Adams–Bohart model. The values of Adams–

Bohart model parameters have been summarized in Table 2. As indicated, the fitting breakthrough curves become worse when C_t/C_0 is higher than 0.5, indicating that the Adams–Bohart model is just appropriate for the initial part of fixed-bed adsorption for removal of Cr (VI) using $\text{Ag}_2\text{O}/\text{SD NC}$. The values of R^2 (from 0.95 to 0.98) indicate the relatively high correlation between the fitting data by a nonlinear method and the experimental data. K_{AB} increases with an increase in initial concentration and bed mass. Conversely, the values of N_0 decrease when there is an increase in bed mass, whereas there is an increase in the value of N_0 with initial concentration and flow rate increasing. These results confirm that the sorption kinetics is surmounted by external mass transfer (Bhaumik *et al.*, 2013; Wang *et al.*, 2015).

3.3.3. Bed depth service time (BDST) model

The lines of $t-z$ at values of C_t/C_0 , 0.2, 0.4 and 0.6 are shown in Figure 9.

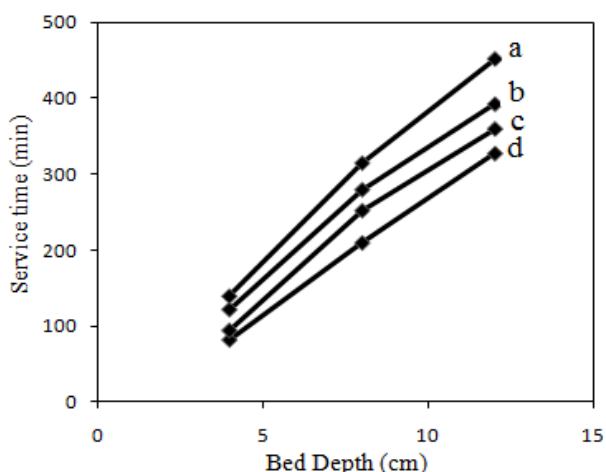


Figure 9. The linear plots of service time (t_b) vs. bed depth at four different breakthrough points ((a) $C_t/C_0=0.05$, (b) $C_t/C_0 = 0.2$, (c) $C_t/C_0 = 0.4$, (d) $C_t/C_0 = 0.6$)

Table 4. BDST equation parameters obtained from Cr (VI) sorption onto $\text{Ag}_2\text{O}/\text{SD NC}$ at different breakthrough points

C_t/C_0	N_0 (mg/L)	K_a (L/mg.min)	Correlation coefficient (R^2)
0.05	1612.41	0.00307	0.9994
0.2	1744.03	0.00167	0.9957
0.4	1776.94	0.00324	0.9908
0.6	2053.35	-0.0168	0.9951

From the values of R^2 , the validity of the BDST model for the present system is confirmed. So, the related constants of BDST according the slopes and intercepts of lines obtained from the linearized form are listed in Table 4. According to Table 4, as the value of C_t/C_0 increased, the rate constant of K_a increased and also the adsorption capacity of the bed per unit bed volume (N_0) increased. Overall, when K_a decreases a progressively with longer C_t/C_0 ratio, this is better and required to avoid breakthrough (Rangabhashiyam *et al.*, 2016c; Cooney, 1999). Pursuant to the regression coefficients ($R^2 > 0.99$),

the match between the experimental data and the model is confirmed.

3.4. Desorption study

In this study, various chemicals were tested to recover the adsorbed Cr (VI) ions from the saturated adsorbent ($\text{Ag}_2\text{O}/\text{SD NC}$). Solution of Cr (VI) with concentration of 25 mg/L was passed through the column with a bed depth of 4 cm and flow rate of 2 mL/min. The results obtained are summarized in Table 5.

Table 5. Desorption of Cr (VI) from the used column applying different washing solutions

Desorption %	Desorption solution			
	Distilled water	NaOH (0.01 M)	HCl (0.01 M)	NaCl (0.01 M)
	< 2	> 98.0	0.00	2.0

As the data show, among the different washing solutions used, it was found that desorption is readily achieved with a dilute solution of NaOH (0.01 M). It was found that the regenerated exhausted column can be used for removal of Cr (VI) without any measurable loss of its sorption capacity. Sorption/desorption processes were repeated at least for three cycles. It was found that more than 98% of previously adsorbed Cr (VI) onto $\text{Ag}_2\text{O}/\text{SD NC}$ can be easily recovered upon treatment with a dilute NaOH solution (0.01M) without any measurable decay in sorption capacity (Rangabhashiyam *et al.*, 2015d). Therefore, $\text{Ag}_2\text{O}/\text{SD NC}$ can be used for removal of highly toxic Cr (VI) ions from the related wastewaters very effectively and efficiently. The exhausted columns can also be simply regenerated for re-use employing a dilute NaOH solution.

4. Conclusion

This study shows that $\text{Ag}_2\text{O}/\text{SD NC}$ is an effective and inexpensive adsorbent for Cr (VI) removal from aqueous solutions. In the fixed-bed column, the adsorption capacity is strongly dependent on the flow rate, pH of feed solution, bed depth and initial feed concentration. The bed volumes slightly increased as flow rate and initial concentration was decreased. As the flow rate was increased, the BTC became steeper, the break point time and adsorbed Cr (VI) ion concentration decreased. With an increase in influent pH, the bed volumes treated decreased. The maximum capacity of $\text{Ag}_2\text{O}/\text{SD NC}$ in column investigation was found to be about 13.4 mg/g under the selected operational conditions (Flow rate= 2 mL/min, $C_i=25$ mg/L, bed depth= 4 cm and pH =2). Thomas, Adams-Bohart and BDST model were used in analysis of column performance and the model parameters were evaluated. The Thomas model was found to be in good agreement with experimental data and could be used for prediction of experimental results as well. The Adams-Bohart model parameters confirm that the sorption kinetics is surmounted by external mass transfer. Also BDST model showed a good correlation of breakthrough curves. The regeneration of exhausted nanocomposite adsorbent after adsorption process can be done with 0.01 M NaOH solution and the regenerated sorbent can be used for adsorption without any loss in its sorption capacity (Reg% > 95).

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Nomenclature

A_c	The area under the breakthrough curve
A	The cross-sectional area of the bed (cm^2)
C_{ad}	The adsorbed concentration (mg/L)
C_0	The concentration of inlet Cr (VI) at time t (mg/L)
C_t	The concentration of effluent Cr (VI) at time t (mg/L)
C_b	The breakthrough Cr (VI) concentration (mg/L)
F	The Flow rate (mL/min)
K_{Th}	The Thomas kinetics model rate constant (mL/min. mg)
K_{AB}	The Bohart–Adams model kinetic constant (L/mg. min)
K_a	The Bed-depth service time model adsorption rate constant (L/mg. min)
m_t	The total amount of Cr (VI) sent to column (g)
N_0	The saturation adsorption capacity per column volume (mg/L)
q_{total}	The total stoichiometric quantity of adsorbed Chromium (VI) (mg)
q_{reg}	The adsorptive capacity of the regenerated column (mg/g)
q_{org}	The sorption capacity of the original column (mg/g)
q_0	The equilibrium Chromium (VI) uptake per g of the adsorbent (mg/g)
S	The removal percentage (%)
t_B	The service time of Bed-depth service time model (min)
t_t	The time equivalent to the total stoichiometric capacity of the column (min)
U_0	The linear flow velocity (cm/min)
V_{eff}	The effluent volume (L)
x	The amount of adsorbent in the column (g)
Z	The bed height of the fix-bed column (cm)